

9th STD SECOND TERM EXAMINATION 2023-24

Chemistry Answer Key

1.	He
2.	b) Statement (i) and (iii) are true.
3.	H ₂ O
4.	Metalloids
5.	c. Loss of electrons.
6.	a. Inner transition elements. b. Lanthanoids
7.	a. 2,8,5 b. 16
8.	a. 2g b. Law of conservation of mass.
9.	a. ¹⁴ Si b. 18
10.	a. X = HSO ₄ ⁻ Y = SO ₄ ²⁻ b. H ₂ SO ₃ or H ₂ CO ₃
11.	a. Sodium hydroxide. b. Red litmus turns to blue colour. Alkali nature.
12.	a. B b. A c. A
13.	a. +2 b. Yes. Oxidation and Reduction take place simultaneously.
14.	a. Mg ²⁺ b. 2,8. c. Solid/ Generally soluble in water/ conducts electricity in aqueous and molten state/ High melting and boiling point.

15.	<p>a. Covalent bond.</p> <p>b. 1</p> <p>c. AlCl_3</p>															
16.	<p>a. Mg</p> <p>b. Cl</p> <p>c. $\text{Cl} + 1\text{e}^- \rightarrow \text{Cl}$</p> <p>d. $\text{Mg} \rightarrow \text{Mg}^{2+} + 2\text{e}^-$</p>															
17.	<p>a. Neutralisation</p> <p>b. NaCl</p> <p>c. Neutral</p> <p>d. The medicine used for reducing acidity in the stomach are known as antacid.</p>															
18.	<p>a. Marble powder and dil.HCl. Surface area increases, the rate of reaction increases.</p> <p>b. $\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2$</p> <p>c. Increase the concentration of reactants.</p>															
19.	<table border="1"> <thead> <tr> <th>A</th> <th>B</th> <th>C</th> </tr> </thead> <tbody> <tr> <td>Lavoiser</td> <td>Metals and Non metals</td> <td>Could not include metalloids</td> </tr> <tr> <td>Doberiner</td> <td>Triads</td> <td>Group of three elements</td> </tr> <tr> <td>Mendeleev</td> <td>Increasing order of atomic mass</td> <td>Prediction of the properties of undiscovered elements</td> </tr> <tr> <td>Moseley</td> <td>Increasing order of atomic number</td> <td>18 groups and 7 periods</td> </tr> </tbody> </table>	A	B	C	Lavoiser	Metals and Non metals	Could not include metalloids	Doberiner	Triads	Group of three elements	Mendeleev	Increasing order of atomic mass	Prediction of the properties of undiscovered elements	Moseley	Increasing order of atomic number	18 groups and 7 periods
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20.	<p>a. 15</p> <p>b. 3</p> <p>c. 8.</p> <p>d. 2, 5</p>															